

TABLE 3.

Composition of the addition-compound formed in SDS/DMLL/CaCl₂ system.

	Na		Ca		C	
	(mol/l)	(%)	(mol/l)	(%)	(mg/l)	(%)
Supernatant solution	1.00×10^{-2}	100	2.11×10^{-2}	84.4	50.8	1.32
Addition-compound	0	0	3.90×10^{-3}	15.6	3.84×10^3	98.7

and the kinds of salt. One should note that as the concentration of CaCl₂ becomes larger than a certain value, the dissolution temperature rapidly increases. This may be attributed to the fact that the addition compound formation between dodecyl sulfate ion and DMLL is caused by the presence of calcium ions.

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ERRATUM

The wrong figure was inadvertently published as Fig. 5 in "Preparation of Geometrical Isomers of Linolenic Acid" by A. Grandgirard, F. Julliard, J. Prevost and J.L. Sebedio in *JAOCS* 64:1434 (October 1987). The correct figure and figure caption are published below.

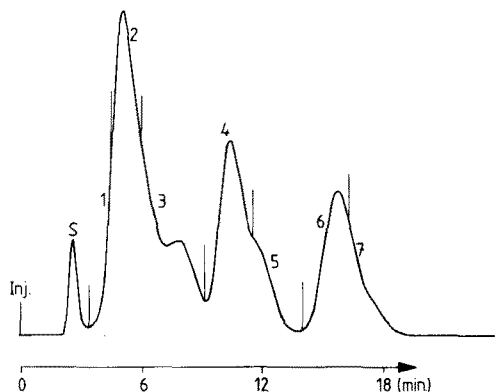


FIG. 5. Preparative HPLC fractionation of a mixture of FAMES isolated from a heated linseed oil (275 C, 12 hr, under nitrogen, same HPLC conditions as in Fig. 2).